

Characterization of 2,3-Dihydroxyterephthalamides as M(IV) Chelators¹

Christine J. Gramer and Kenneth N. Raymond*

Chemical Sciences Division, Lawrence Berkeley National Laboratory, and Department of Chemistry, University of California, Berkeley, California 94720-1460

Received March 19, 2004

The ligand *N,N'*-diethyl-2,3-dihydroxyterephthalamide (ETAM) has been characterized as a chelator for Zr(IV), Ce(IV), and Th(IV). The K⁺ salts of the complexes [Zr(ETAM)₄]⁴⁻, [Ce(ETAM)₄]⁴⁻, and [Th(ETAM)₄]⁴⁻ were prepared in a MeOH solution containing H₂ETAM, the corresponding M(acac)₄, and 4 equiv of KOH. Single-crystal X-ray diffraction analyses are reported for K₄[Zr(ETAM)₄] (C2/c, Z = 8, *a* = 27.576(3) Å, *b* = 29.345(3) Å, *c* = 15.266(2) Å, α = 90°, β = 118.688(4)°, γ = 90°), [Me₃BnN]₄[Th(ETAM)₄] (P1̄, Z = 2, *a* = 13.7570(3) Å, *b* = 13.9293(3) Å, *c* = 26.9124(6) Å, α = 99.941(1)°, β = 94.972(1)°, γ = 103.160(1)°), and the dimeric (NMe₄)₄[Th(ETAM)₃MeOH]₂ (P2₁/c, Z = 4, *a* = 18.2603(9) Å, *b* = 18.5002(9) Å, *c* = 19.675(1) Å, β = 117.298(1)°). Solution thermodynamic studies were used to determine formation constants (log *K*_f and esd) for Th(IV)–ETAM log *K*₁₁₀ = 17.47(1), log *K*₁₂₀ = 13.23(1), log *K*₁₃₀ = 8.28(3), log *K*₁₄₀ = 6.57(6), and log β₁₄₀ = 45.54(5). These results support the hypothesis that the terephthalamides are high-affinity chelators for the actinide(IV) ions and thus promising ligands for use in nuclear waste remediation.

Introduction

The purification of radioactive fission materials and spent fuels from the nuclear industry has resulted in large volumes of highly radioactive waste that must ultimately be safely stored in a geological repository. Reducing the volume of this waste is a major goal of nuclear waste remediation, and this can be aided through the selective removal of actinide metals from the complex waste mixtures that contain a variety of organic and inorganic components. A biomimetic approach based on the observation that Fe(III) and Pu(IV) are both hard metal ions with the same charge-to-size ratio and a preference for oxygen donor ligands has successfully been applied to the development of actinide sequestering agents.² Siderophores are powerful bacterial chelators of iron, and many of the most potent of these chelators incorporate catecholamide (CAM) ligand groups (Figure 1).³ The extraordinary Fe(III) chelating properties of the CAMs prompted the development of other catecholate-based ligands contain-

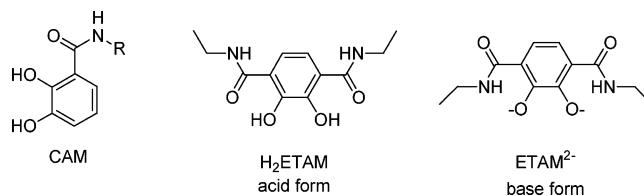


Figure 1. Structure of the catecholamide ligands and *N,N'*-diethyl-2,3-dihydroxyterephthalamide (H₂ETAM) in its protonated and deprotonated forms.

ing additional carboxyl or sulfate groups and the characterization of these ligands as powerful chelators for Fe(III) and An(IV) ions confirmed the validity of this biomimetic approach to M(IV) chelation.^{4–6}

Early studies with the catechol dianion and the eight-coordinate, tetravalent metals U(IV), Ce(IV), Th(IV), and Hf(IV) indicated that catechol itself is a strong chelator of high oxidation state metals.^{7,8} A derivative of the naturally occurring catecholamide which contains a second amide

* Author to whom correspondence should be addressed. E-mail: raymond@socrates.berkeley.edu.

- (1) This paper is no. 52 in the series "Specific Sequestering Agents for the Actinides." For the previous paper, see: Guilmette, R. A.; Hakimi, R.; Durbin, P. W.; Xu, J.; Raymond, K. N. *Radiat. Prot. Dosim.* **2003**, *105*, 527–534.
- (2) Gorden, A. E. V.; Xu, J.; Raymond, K. N.; Durbin, P. W. *Chem. Rev.* **2003**, *103*, 4207–4282.
- (3) Raymond, K. N.; Durbin, P. W. In *The First Hanford Separation Science Workshop Publication PNL-SA-21775*, Richmond, WA, 1993.

- (4) Kappel, M. J.; Nitsche, H.; Raymond, K. N. *Inorg. Chem.* **1985**, *24*, 605–611.
- (5) Garrett, T. M.; Miller, P. W.; Raymond, K. N. *Inorg. Chem.* **1989**, *28*, 128–133.
- (6) Whisenhunt, D. W., Jr.; Neu, M. P.; Hou, Z.; Xu, J.; Hoffman, D. C.; Raymond, K. N. *Inorg. Chem.* **1996**, *35*, 4128–4136.
- (7) Sofen, S. R.; Abu-Dari, K.; Freyberg, D. P.; Raymond, K. N. *J. Am. Chem. Soc.* **1978**, *100*, 7882–7887.
- (8) Sofen, S. R.; Cooper, S. R.; Raymond, K. N. *Inorg. Chem.* **1979**, *18*, 1611–1616.

group, the 2,3-dihydroxyterephthalamide (TAM), was one of these ligands characterized as an Fe(III) chelator.⁵ In the progression from catechol, to catecholamide (CAM), to terephthalamide (TAM), the addition of electron-withdrawing amides lowers the protonation constants of the phenolic oxygens but does not decrease the effective strength of Fe(III) binding. Thus, the TAMs are also a logical candidate for actinide(IV) chelation. Previously, we have reported the syntheses and incorporation of these ligands into water-soluble polymers for the extraction of Pu(IV).^{9,10} Here, we present the coordination chemistry of a simple bidentate alkyl TAM ligand, *N,N'*-diethyl-2,3-dihydroxyterephthalamide (ETAM), with the tetravalent metals Zr(IV), Ce(IV), and Th(IV) as models for Pu(IV).

The size of the metals increases in the order Zr(IV) < Ce(IV) ~ Pu(IV) < Th(IV);¹¹ thus, Ce(IV) is a convenient Pu(IV) analogue since the ionic radii of Ce(IV) and Pu(IV) are similar^{11,12} and because the aqueous coordination chemistry of Ce(IV) parallels that of Pu(IV), giving identical shifts in the M(IV)/M(III) redox potentials.⁴ Zr(IV) and Th(IV) are useful since neither is electrochemically active. The similar ionic radii, coordination number, and preferred oxidation state of Th(IV) make it a good analogue for the more radioactive and hazardous Pu(IV),^{13,14} while the smaller Zr(IV) allows for comparisons to determine the structural effect of the 5f electrons. We report the synthesis of these M(IV) ETAM complexes, crystallographic determination of the Zr(IV) and Th(IV) ETAM structures, and the stability constants for the Th(IV) ETAM complex.

Experimental Section

General Methods. All reagents and solvents were purchased from either Aldrich Chemical Co. or Fisher Scientific and used as purchased. Solvents were dried over activated alumina and stored over 4 Å molecular sieves. All reactions were carried out under Ar, and all solvents were degassed by evacuating the flask and filling with argon 3 times. Water was distilled and further purified by a Millipore cartridge system (resistivity $18 \times 10^6 \Omega$). ¹H and ¹³C NMR spectra were obtained on a Bruker DRX500 (500 MHz) or Bruker AMX400 spectrometer (400 MHz) as noted. All NMR samples were taken in DMSO-*d*₆. Microanalyses and mass spectra were performed by the Microanalytical Services Laboratory in the College of Chemistry, University of California, Berkeley, CA. H₂ETAM and Ce(acac)₄ were synthesized by published methods.^{5,9,15}

In each of the following cases, the ligand H₂ETAM (100 mg, 0.400 mmol) was dissolved in 10 mL of degassed MeOH. The metal acetylacetonate (M(acac), 0.100 mmol) was dissolved in 5 mL of MeOH, filtered through a cotton plug to remove undissolved particles, and added to the ligand solution. The reaction was heated

to reflux temperature for 16 h and then concentrated by evaporation to a tan solid. This was redissolved in MeOH and degassed, and 0.5 M KOH (0.8 mL, 0.4 mmol) was added. After the solution was stirred for 3 h, the solvent was evaporated, and the tan solid was dissolved in 0.5 mL MeOH and purified by means of a Sephadex LH-20 column. The desired fraction was collected and evaporated to a tan solid.

K₄[Zr(ETAM)₄]. Yield: 92 mg, 74%. ¹H NMR (DMSO-*d*₆): δ 0.83 (t, 3H), 3.00 (q, 2H), 6.53 (s, 2H), 10.88 (t, NH). Anal. Calcd for K₄Zr(ETAM)₄·17H₂O: C, 37.08; H, 5.83; N, 7.21. Found: C, 37.17; H, 5.98; N 7.04. MS [ESI-] (*m/z*): 420.3, [ZrL₃]²⁻. Crystals were grown from diffusion of diethyl ether into 2 mL of ethanol containing 4 mg of the complex.

K₄[Ce(ETAM)₄]. A black solid formed. Yield: 92 mg, 71% yield. ¹H NMR (DMSO-*d*₆): δ 0.91 (t, 3H), 3.08 (q, 2H), 6.56 (s, 2H), 10.69 (t, NH). Anal. Calcd for K₄Ce(ETAM)₄·5H₂O: C, 41.55; H, 4.79; N, 7.94. Found: C, 41.49; H, 4.65; N, 8.08.

K₄[Th(ETAM)₄]. A tan solid formed. Yield: 100 mg, 72% yield. ¹H NMR (DMSO-*d*₆): δ 0.91 (t, 3H), 3.09 (q, 2H), 6.63 (s, 2H), 10.87 (t, NH). Anal. Calcd for K₄Th(ETAM)₄·11H₂O: C, 36.31; H, 4.95; N, 7.06. Found: C, 36.38; H 4.73; N 6.90. FAB-MS(-) (*m/z*): 1349.2, [ThL₄·3K]⁻. Crystals of the [Th(ETAM)₄]⁴⁻ complex were prepared as the tetrakis(trimethylbenzylammonium) salt by using a stoichiometric amount of [Me₃BnN]OH as the base instead of KOH during the metal complex synthesis. Crystals were grown from diffusion of either methyl *tert*-butyl ether or diethyl ether into 1 mL of methanol.

(Me₄N)₄[Th(ETAM)₃]₂. H₂ETAM (76 mg, 0.30 mmol) was dissolved in 10 mL of degassed MeOH. Th(acac)₄ (62 mg, 0.10 mmol) was dissolved in 5 mL of MeOH and added with filtering to the ligand solution. The reaction was heated to reflux temperature for 16 h and then concentrated by means of evaporation to a tan solid. This was redissolved in MeOH and degassed, and 0.5 M KOH (0.6 mL, 0.3 mmol) was added. After the solution was stirred for 3 h, the solvent was evaporated, and the tan solid was dissolved in 0.5 mL MeOH and applied to a Sephadex LH-20 column. The desired fraction was collected and evaporated to a tan solid (94 mg, 86% yield). Anal. Calcd (found) for K₄Th₂(ETAM)₆·16H₂O: C, 35.88 (35.38); H, 4.85 (4.59); N, 6.97 (6.76). FAB-MS(-) (*m/z*): 1021, [Th₂L₆·2K]²⁻. Crystals were grown from diffusion of either methyl *tert*-butyl ether or diethyl ether into 1 mL of methanol containing 4 mg of the K⁺ salt of the complex and 4 mg of NMe₄Cl.

Structure Determination and Refinement. All X-ray structure data sets were collected on a Siemens SMART area detector diffractometer.¹⁶ Crystals were mounted on quartz capillaries in Paratone oil and were cooled in a nitrogen stream on the diffractometer. Peak integrations were performed using the Siemens SAINT software package.¹⁷ Space group determinations were done using the software XPREP. The two Th structures were solved and refined using the teXsan software package,¹⁸ and the Zr structure was solved and refined using SHELXTL.¹⁹ All hydrogen atoms were fixed at calculated positions and their thermal parameters refined isotropically. For [Zr(ETAM)₄]⁴⁻, only the Zr, oxygen, and nitrogen atoms of the ligand were refined anisotropically. For [Th(ETAM)₄]⁴⁻ and the [Th(ETAM)₃]₂⁴⁻ dimer, all non-hydrogen

(9) Gramer, C. J.; Raymond, K. N. *Org. Lett.* **2001**, *3*, 2827–2830.
 (10) Gramer, C. J.; Raymond, K. N.; Jarvinen, G. D.; Schroeder, N. C.; Robison, T. W.; Smith, B. F. *Sep. Sci. Technol.* **2004**, *39*, 321–339.
 (11) Shannon, R. D. *Acta Crystallogr.* **1976**, *A32*, 751–767.
 (12) Shannon, R. D.; Prewitt, C. T. *Acta Crystallogr., Sect. B* **1969**, *B25*, 925.
 (13) Aspinall, H. C. *Chemistry of the f-Block Elements*; Gordon and Breach Publishers: Amsterdam, 2001.
 (14) Choppin, G. R.; Nash, K. L. *Radiochim. Acta* **1995**, *70/71*, 225–236.
 (15) Xu, J.; Radkov, E.; Ziegler, M.; Raymond, K. N. *Inorg. Chem.* **2000**, *39*, 4156–4164.

(16) SMART, Area-Detector Software Package; Siemens Industrial Automation Inc.: Madison, WI, 1994.
 (17) SAINT, SAX Area-Detector Integration Program, 4.024 ed.; Siemens Industrial Automation Inc.: Madison, WI, 1994.
 (18) teXsan, Crystal Structure Analysis Package; Molecular Structure Corp.: The Woodlands, TX, 1992.
 (19) SHELXTL, Crystal Structure Analysis Determination Package; Siemens Industrial Automation Inc.: Madison, WI, 1994.

Table 1. Crystal Structure Details for $K_4[Zr(ETAM)_4]$, $[Me_3BnN]_4[Th(ETAM)_4]$, and $[NMe_4]_4[Th(ETAM)_3MeOH]_2$

	$K_4[Zr(ETAM)_4]$	$[Me_3BnN]_4[Th(ETAM)_4]$	$[NMe_4]_4[Th(ETAM)_3MeOH]_2$
fw	1336.64	1834.02	1529.63
temp, °C	−114	−124	−117
cryst syst	monoclinic	triclinic	monoclinic
space group	$C2/c$	$P\bar{1}$	$P2_1/c$
unit cell dimens			
<i>a</i> , Å	27.576(3)	13.7570(3)	18.2603(9)
<i>α</i> , deg	90	99.941(1)	90
<i>b</i> , Å	29.345(3)	13.9293(3)	18.5002(9)
<i>β</i> , deg	118.688(4)	94.972(1)	117.298(1)
<i>c</i> , Å	15.266(2)	26.9124(6)	19.675(1)
<i>γ</i> , deg	90	103.160(1)	90
<i>V</i> , Å ³ ; <i>Z</i>	10837.0(2); 8	4903.4(2); 2	5906.3(5); 4
<i>d</i> _{calcd} , g cm ^{−3}	1.638	1.242	1.72
cryst size, mm	0.27 × 0.40 × 0.15	0.23 × 0.17 × 0.12	0.16 × 0.10 × 0.09
abs corr <i>μ</i> , cm ^{−1}	5.92	15.9	26.25
reflens colld	17 366	25 460	24 261
indep reflens	7210 [R(int) = 0.1371]	15 569 [R(int) = 0.044]	8795 [R(int) = 0.074]
data/param	2605 /364	11 123/874	4195/585
goodness-of-fit ^{a,b}	1.016 ^a	1.58 ^b	1.04 ^b
final R indices ^{a,b}	R1 = 0.1110 ^a wR2 = 0.2969	R1 = 0.052 ^b R _w = 0.060	R1 = 0.032 ^b R _w = 0.030

^a Solved using SHELXTL;¹⁹ goodness-of-fit on F^2 ; final R indices [$I > 2\sigma(I)$]. ^b Solved using TEXAN;¹⁸ goodness-of-fit on F ; final R indices [$I > 3\sigma(I)$].

atoms were refined anisotropically. The crystallographic details of each structure are given in Table 1.

Solution Thermodynamics. Equilibrium constants were measured at 25 °C in aqueous solution with a supporting electrolyte of 0.1 M KCl for Th(IV). They were determined as cumulative formation constants (β_{mlh} , where M, L, and H refer to the metal, ligand, and proton, respectively):

$$\beta_{mlh} = \frac{[M_m L_l H_h]}{[M]^m [L]^l [H]^h}$$

For convenience, these are discussed as stepwise association constants, either for complex formation (K_{ln0}) or ligand protonation (K_{0lh}). For example,

$$K_{ln0} = \frac{[ML_n]}{[ML_{n-1}][L]} = \frac{\beta_{ln0}}{\beta_{l(n-1)0}} \quad K_{0lh} = \frac{[LH_n]}{[LH_{n-1}][H]} = \frac{\beta_{0lh}}{\beta_{0l(h-1)}}$$

Apparatus. A temperature-controlled 100 mL titration flask with attached quartz UV–vis cell has been described in detail elsewhere.²⁰ An Accumet pH-meter (model AR15 or 15) and a Corning glass-bulb combination electrode were used for electrode potential measurements. Metrohm autoburets (Dosimat 665/765) were used for titrant dosing. For the collection of absorbance data, a Hewlett-Packard 8452a spectrophotometer (diode array) was used. The lamp was always turned on for at least 1 h before use to allow the lamp to come to operating temperature, and all absorbance measurements were less than 1.1 absorbance units. The autoburets, pH meters, and spectrophotometer were controlled by a personal computer using modules of the LABVIEW²¹ programming environment, allowing automated titration execution. Acid and base solutions (0.1 M HCl and KOH) were prepared using J.T. Baker Dilut-it ampules and standardized to $\pm 0.2\%$. An analytical grade balance (± 0.05 mg) was used.

Standardization of Th(IV) Stock Solution.²² A Th(IV) solution was prepared by dissolving ThCl₄ in HCl and HNO₃ such that the

[Th(IV)] was approximately 5 mM and the [H⁺] was 0.2 M. In a 40 °C temperature-controlled cell, 2.5 mL of the Th(IV) solution was diluted into 50 mL of doubly distilled H₂O and 4 drops of a 0.1% pyrocatechol violet solution were added. The solution was titrated with 0.1018 M EDTA until the pink color disappeared. This was repeated 5 times, and a final concentration of 0.0506(2) M Th(IV) was calculated for the stock solution.

Electrode Calibration. The standard electrode calibration was conducted before each titration by following published protocol.²⁰ The electrode was calibrated by adding 2.000 mL of HCl (ca. 0.1 M) to 50.0 mL of 0.1 M KCl, and the solution was titrated with KOH (ca. 0.1 M) to pH 11.6. The data were analyzed using the program GLEE²³ allowing refinement of E° and slope. For low-pH titrations, a correction was made for the junction potential.²⁰

Th(IV) ETAM Titrations. Into 50 mL of 0.1 M KCl was added an aliquot of 0.024089 M H₂ETAM solution in MeOH (0.05189 mL, 0.00125 mmol), an aliquot of a 0.0506 M Th(IV) stock solution (0.0049 mL, 0.00025 mmol), HEPES (4.44 mg, 0.019 mmol), MES (4.7 mg, 0.022 mmol), NH₄Cl (1.6 mg, 0.03 mmol), and 1.000 mL of 0.0988 M HCl. This was titrated with 0.1003 M KOH from pH 2.8 to 10.5 and then back to pH 2.8 with 0.0988 M HCl, with a constant pH increment of 0.05 such that 67–82 data points, each consisting of pH and absorbance spectrum, were collected in each of the forward and reverse titrations. This was repeated 3 times, for a total of 6 titrations. An equilibration time of 3 min after addition of the titrant was used. The specifics of data collection have been detailed elsewhere.²⁰ All absorbance measurements were less than 1.1 absorbance units. The spectrum was recorded, and the data points from 320 to 420 nm were used in the data processing. All data was analyzed using the program pHAB.²⁴ The following values of log β for formation of metal hydroxides were included in the refinement: [Th(OH)]³⁺, −3.8; [Th(OH)₂]²⁺, −7.66; [Th(OH)₃]⁺, −12.76.²⁵ The p*K*_a values for H₂ETAM are 6.1 and 11.0 and were previously determined.⁵

A separate determination of the first equilibrium constant for the metal–ligand complex was conducted by a titration from

(20) Johnson, A. R.; O'Sullivan, B.; Raymond, K. N. *Inorg. Chem.* **2000**, *39*, 2652–2660.

(21) LABVIEW, 5.0.1 ed.; National Instruments Corp.: Austin, TX.

(22) Schwarzenbach, G.; Flaschka, H. *Complexometric Titrations*; Methuen and Co., Ltd.: London, 1969.

(23) Gans, P.; O'Sullivan, B. *Talanta* **2000**, *51*, 33–37.

(24) Gans, P.; Sabatini, A.; Vacca, A. *Ann. Chim.* **1999**, *89*, 45–49.

(25) Baes, C. F.; Mesmer, R. E. *The Hydrolysis of Cations*; Wiley: New York, 1986.

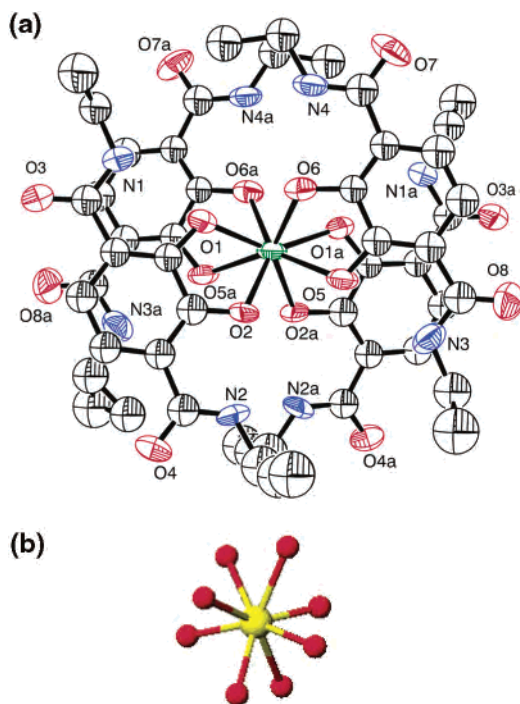


Figure 2. (a) ORTEP diagram of the $[\text{Zr}(\text{ETAM})_4]^{4-}$ complex (50% probability) showing the square antiprismatic coordination geometry, orientation of the amides for hydrogen-bonding, and numbering scheme for the nitrogen and oxygen atoms. (b) Details of the coordination geometry of the Zr(IV) ion in the $[\text{Zr}(\text{ETAM})_4]^{4-}$ complex.

pH 3.5 to 2.9 with corrections to the pH measurement for variation in electrode junction potential as stated above. An aliquot of a 0.0241 mM H_2ETAM solution in MeOH (0.5189 mL, 0.0125 mmol, 1% MeOH total) and an aliquot of a 0.0506 M Th(IV) stock solution (0.049 mL, 0.0025 mmol) were added to 50 mL of 0.1 M KCl, and the solution was titrated from pH 3.5 to 2.9 with 0.0998 M HCl at constant pH increments such that 25 data points were collected.

Results and Discussion

Synthesis and X-ray Crystal Structures of M(IV) ETAM Complexes. Three complexes with composition $\text{K}_4[\text{M}(\text{ETAM})_4]$ ($\text{M} = \text{Zr}(\text{IV}), \text{Ce}(\text{IV}), \text{Th}(\text{IV})$) were prepared by a ligand exchange reaction in which the appropriate $\text{M}(\text{acac})_4$ and 4 equiv of H_2ETAM were heated in refluxing MeOH for 16 h and then treated with 4 equiv of methanolic KOH to fully deprotonate the complex. After evaporation, the complex was dissolved in MeOH and separated from impurities by Sephadex LH-20 column chromatography.

The structure of $\text{K}_4[\text{Zr}(\text{ETAM})_4]$ shown in Figure 2 illustrates the square antiprismatic coordination geometry. As seen in previously characterized metal-TAM complexes, the amide is oriented such that the amide proton forms a hydrogen bond with the catechololate oxygen.²⁶ The approach of the ligands to the Zr(IV) center forms an angle of $\sim 115^\circ$ between the catechololate plane, phenolic oxygens, and Zr(IV) with the Zr(IV) lying 0.76 or 0.87 Å out of the plane of the catechol ring. In the view shown, the planes of the two ligands on top intersect each other with an angle of 147° ,

and the two catechol planes on the left intersect each other with an angle of 25° . The Zr–O bond lengths range from 2.14 to 2.27 Å. One of the K^+ counterions sits in the cleft above the four catechololate oxygens and is coordinated by each oxygen. The remaining K^+ counterions are coordinated by amide oxygens from neighboring complexes (not shown) and in this way link all of the $[\text{Zr}(\text{ETAM})_4]^{4-}$, forming large voids in the extended structure that accommodate disordered solvent.

The quantitative shape measure has been introduced to compare the eight-coordinate polyhedra of several Ce(IV) hydroxypyridinone (HOPO) crystal structures with idealized eight-coordinate polyhedra.¹⁵ The three idealized polyhedra for eight-coordinate metal complexes are square antiprism (D_{4d}), bicapped trigonal prism (C_{2v}), and trigonal dodecahedron (D_{2d}). For the shape measure (S), the dihedral angles of adjacent planes from crystal structure coordinates are calculated and compared to the dihedral angles of the D_{4d} , D_{2d} , and C_{2v} geometries:

$$S = \min \left[\frac{1}{m} \sqrt{\sum_{i=1}^m (\delta_i - \theta_i)^2} \right]$$

Here m = number of edges, δ = angle between normals of adjacent faces, δ_i = observed dihedral angle along the i th edge of δ , and θ = same angle of corresponding ideal polytopal shape θ . The smallest S value is the one closest to describing the coordination geometry. In the case of $[\text{Zr}(\text{ETAM})_4]^{4-}$ the following shape measures were calculated: $S(D_{4d}) = 12.8^\circ$; $S(C_{2v}) = 16.9^\circ$; $S(D_{2d}) = 17.3^\circ$. This indicates that the geometry most closely resembles a square antiprism, as is clear from the ORTEP diagrams.

Crystals of the K^+ salt of the analogous Th(IV) complex of a quality suitable for characterization by XRD were never obtained even from a variety of crystallization methods in various solvent, so the $[\text{Th}(\text{ETAM})_4]^{4-}$ as the tetrakis-(trimethylbenzylammonium) salt was prepared using a stoichiometric amount of $[\text{Me}_3\text{BnN}]\text{OH}$ as the base instead of KOH during the metal complex synthesis. This structure, shown in Figure 3, is also eight-coordinate, although the structure has a coordination geometry different from that of the Zr structure and the angle between the catechol ring, phenolic oxygens, and Th(IV) is 180° . The shape measure (S) was used to analyze the coordination geometry, which is not as obvious in this structure. This provided values of $S(D_{4d}) = 22.6^\circ$, $S(C_{2v}) = 18.6^\circ$, and $S(D_{2d}) = 15.6^\circ$, indicating that the coordination geometry around the Th(IV) lies between bicapped trigonal prismatic (C_{2v}) and trigonal dodecahedron (D_{2d}) but closer to D_{2d} . The Th(IV)–oxygen bond lengths are between 2.38 and 2.49 Å.

These Zr(IV) and Th(IV) ETAM structures have coordination geometries which differ from each other and those found in the four crystal structures of $[\text{M}(\text{catecholato})_4]^{4-}$, where $\text{M} = \text{U}(\text{IV}), \text{Th}(\text{IV}), \text{Ce}(\text{IV}),$ and $\text{Hf}(\text{IV})$.^{7,8} The catechololate complexes all have D_{2d} (trigonal dodecahedron) symmetry and all have identical unit cell contents. Th–oxygen bond lengths in the Th catechololate structure of 2.4 Å are close to

(26) Garrett, T. M.; Cass, M. E.; Raymond, K. N. *J. Coord. Chem.* **1992**, *25*, 241–253.

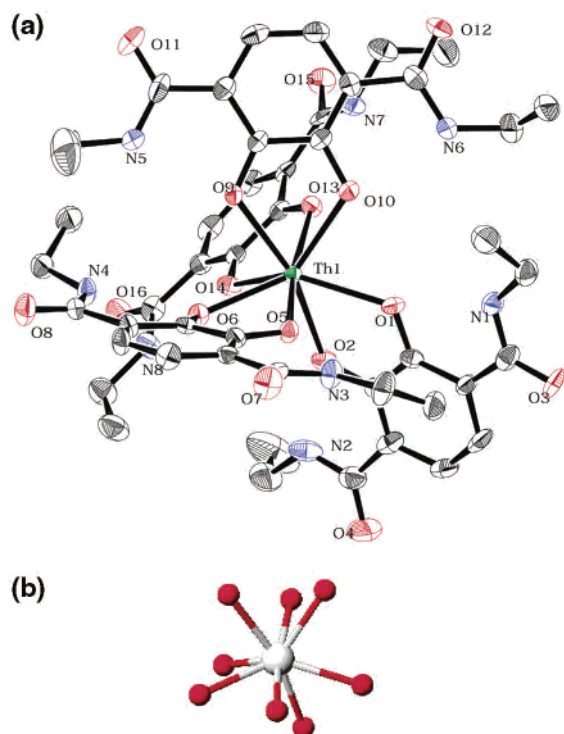


Figure 3. (a) ORTEP diagram of the $[\text{Th}(\text{ETAM})_4]^{4-}$ complex (50% probability) and numbering scheme. (b) Details of the coordination geometry of the Th(IV) ion in the $[\text{Th}(\text{ETAM})_4]^{4-}$ complex.

those of the Th TAM structure. The Hf–oxygen bond lengths of 2.2 Å are close to those of the Zr TAM complex (Hf(IV) and Zr(IV) are the same size). In the $[\text{U}(\text{catecholato})_4]^{4-}$ structure, a small difference was seen in the U(IV)–oxygen bond lengths; however, this distortion was not present in the Th(IV) or Ce(IV) catecholate structures. This difference in U–oxygen bond lengths was attributed to a ligand field effect from the $5f^2$ electrons. The Hf(IV) structure also had a small distortion, but this was attributed to the small size of Hf(IV) and the close contacts between adjacent catechol ligands. Th(IV) and Hf(IV) are f^0 and d^0 metals so there is no electronic preference for a particular coordination geometry. In the TAM structures presented here, the difference is attributed to the different counterions: the K^+ ion coordinates to the catecholate oxygens in the Zr(IV) structure and distorts the ligands, which in turn affects the geometry around the metal.

If the Th ETAM complex is instead prepared with a 3:1 ratio of H_2ETAM –Th(IV), light tan cube-shaped crystals readily form after diffusion of diethyl ether into a methanol solution of the complex. The complex crystallizes in $P2_1/c$, shown in Figure 4, as a dimer with the formula $[\text{NMe}_4]_4[\text{Th}(\text{ETAM})_3\text{MeOH}]_2$. Four dimers are present in the unit cell, and the dimer sits on a special position with inversion symmetry. The coordination environment around the Th(IV) ion consists of two TAM units bound in the expected catecholate fashion. The third TAM also binds Th(IV) with its catecholate oxygens; however, one of the catecholate oxygens bridges to the second Th(IV) center. The amide *ortho* to this phenolic oxygen is oriented such that the carbonyl oxygen also binds to the second Th(IV) in a

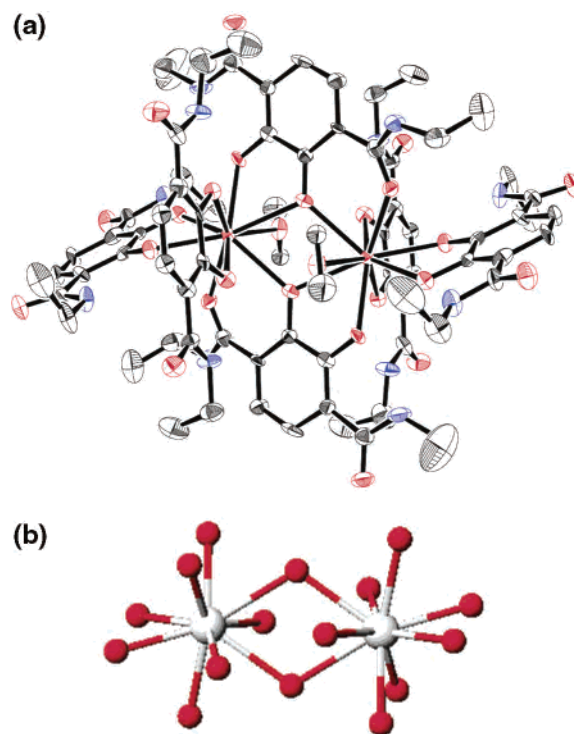


Figure 4. (a) ORTEP diagram of $[\text{NMe}_4]_4[\text{Th}(\text{ETAM})_3\text{MeOH}]_2$ dimeric complex (50% probability). (b) Details of the coordination geometry around the Th(IV) ions in the $[\text{NMe}_4]_4[\text{Th}(\text{ETAM})_3\text{MeOH}]_2$ complex.

Table 2. Formation Constants for Th(IV) ETAM Complexes

K_{mlh}	$\log K_{mlh}$	K_{mlh}	$\log K_{mlh}$
K_{110}	17.47(1)	K_{140}	6.57(6)
K_{120}	13.23(1)	β_{140}	45.54(5)
K_{130}	8.28(3)		

salicylate mode (detail shown in Figure 4b). In all other TAM and CAM metal complex structures, the amide is oriented with the oxygen pointing away from the metal such that the amide proton can hydrogen bond to the deprotonated, coordinating catecholate oxygen. In total, each Th(IV) is coordinated by five phenolic oxygens, two bridging phenolic oxygens, one amide oxygen, and one MeOH, for a total coordination number of 9, consistent with the larger size of Th(IV) as compared to Zr(IV) (ionic radius of 1.01 Å vs 0.86 Å).¹¹

Stability Constant Determination for Th(IV) ETAM Complex. The formation constants for the Th ETAM complex were determined by acid and base titrations of a solution containing Th(IV) and H_2ETAM , and the constants are summarized in Table 2. The pH-dependent sequential addition of ETAM^{2-} ligands to the Th(IV) center was monitored by the distinct changes in the UV–vis spectrum between 250 and 280 nm and 340 and 375 nm attributed to the π to π^* transition of the ligand (Figure 5).⁵ The concentration of Th(IV) and H_2ETAM were carefully chosen to avoid precipitation of the neutral $\text{Th}(\text{ETAM})_2$ during the titration while still maintaining sufficient intensity in absorbance values. Due to these low concentrations, not enough buffering is afforded by H_2ETAM so three noncoordinating buffers, MES, HEPES, and NH_4Cl , were added at ~ 0.4 mM to ensure sufficient buffering. A second “low pH” titration

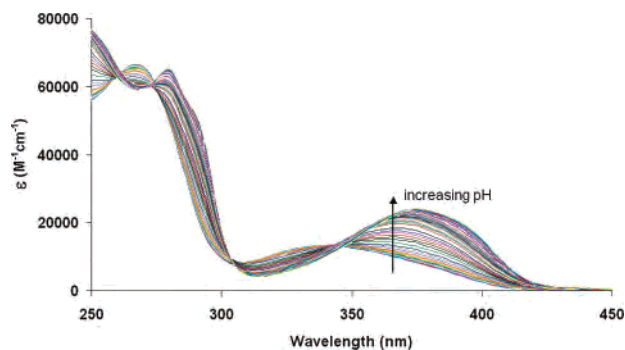


Figure 5. Representative spectra from the Th(IV) ETAM UV-vis spectrophotometric titration.

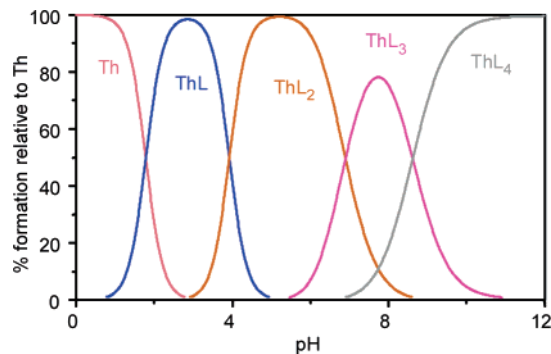


Figure 6. Speciation diagram for the Th(IV)-ETAM system calculated with $[\text{Th(IV)}] = 1 \mu\text{M}$ and $[\text{ETAM}^{2-}] = 10 \mu\text{M}$.

from pH 3.1 to 1.5 was conducted to confirm K_{110} , and in this case the disappearance of the ThL^{2+} complex was monitored spectrophotometrically. The data fit very well to a simple model of sequential addition of ligand to the metal center. Inclusion of mixed hydroxo-ligand-Th(IV) species or dimers did not agree at all with the observed data. H_2ETAM is a high-affinity chelator for Th(IV) with an overall formation constant $\log \beta_{140}$ of 45.54. The speciation diagram shown in Figure 6 (calculated with 10^{-5} M H_2ETAM and 10^{-6} M Th) illustrates how the composition of the solution changes with pH showing that the complex is not fully formed until pH 10 but starts forming at pH 1.

The relative strength of ETAM as a chelator of Th(IV) can also be evaluated by calculating pM. The pM value is defined as the negative logarithm of the free metal concentration and can be calculated for any pH, concentration of metal, and concentration of ligand.²⁷ While the standard concentrations for a pM calculation are $1 \mu\text{M}$ metal and $10 \mu\text{M}$ ligand at a pH of 7.4, in this case it is more useful to calculate pM over the entire pH range since waste samples can vary widely. Shown in Figure 7 is a plot of pM vs pH for the Th(IV) ETAM system, calculated with metal and ligand concentrations of 1 and $10 \mu\text{M}$, illustrating the amount of actinide that would be expected to remain in a waste stream after treatment with a TAM.

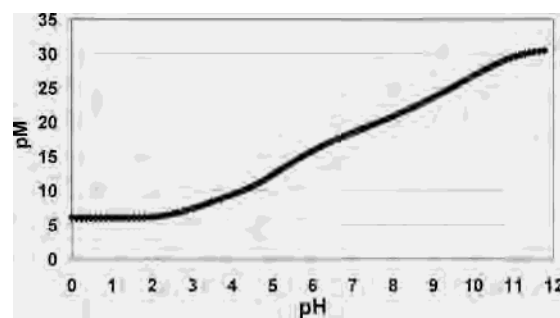


Figure 7. Graph of pM vs pH for Th ETAM calculated with $[\text{M(IV)}] = 1 \mu\text{M}$ and $[\text{ETAM}^{2-}] = 10 \mu\text{M}$.

Conclusion

Understanding the basic coordination chemistry of the TAMs with actinide cations is crucial for determining their potential utility for us as actinide chelators in nuclear waste remediation. The simple bidentate 2,3-dihydroxyterephthalamide ligand ETAM has been evaluated in solution and the solid state as a M(IV) chelator. The complexes of Zr(IV), Ce(IV), and Th(IV) are straightforward to synthesize through a ligand exchange reaction, providing useful models for Pu(IV). The crystal structures show that these complexes can adopt different coordination geometries and coordination number. In the absence of a sufficient amount of ligand to saturate the coordination sphere of Th(IV), the hydrogen bond between the amide proton and catecholate oxygen can be broken to allow the ligand to coordinate through the amide carbonyl oxygen. Solution thermodynamic studies reveal that the TAMs have very high binding constants with Th(IV) and therefore should be well suited for chelating Pu(IV) in a variety of waste streams. With the appropriate choice of amide substituents, the TAMs should be well suited as liquid-liquid extractants for Pu(IV) or incorporation onto a solid support.

Acknowledgment. The authors acknowledge the assistance of Dr. Darren Johnson with the X-ray crystallography, Dr. Brendon O'Sullivan with the solution thermodynamics, and Dr. Jide Xu with synthetic advice. We also thank Dr. Anne E. V. Gorden for the preparation of this manuscript. This research was supported by the Director, Office of Energy Research, Office of Basic Energy Sciences, Chemical Sciences Division, U.S. Department of Energy (DOE), under Contract No. DE-AC03-76SF00098. Additionally, this research was supported in part under Grant No. SF17SP23, Environmental Management Science Program, Office of Science and Technology, Office of Environmental Management, U.S. DOE.

Supporting Information Available: CIF files with the details for the datasets from the crystals of ZrETAM, ThETAM, and the dimeric Th_2ETAM_6 . This material is available free of charge via the Internet at <http://pubs.acs.org>.

(27) Weitzl, F. L.; Raymond, K. N. *J. Am. Chem. Soc.* **1980**, *102*, 2289-2293.